30-10 Decay Series

It is often the case that one radioactive isotope decays to another isotope that is also radioactive. Sometimes this daughter decays to yet a third isotope which also is radioactive. Such successive decays are said to form a decay series. An important example is illustrated in Fig. 30-11. As can be seen, ²³⁸U decays by α emission to $^{234}_{90}$ Th, which in turn decays by β decay to $^{234}_{91}$ Pa. The series continues as shown, with several possible branches near the bottom, ending at the stable lead isotope, ²⁰⁶₈₂Pb. The two last decays can be

$$^{206}_{81}\text{Tl} \rightarrow ^{206}_{82}\text{Pb} + e^- + \bar{\nu},$$

or

$$^{210}_{84}Po \rightarrow ^{206}_{82}Pb + \alpha$$
.

Other radioactive series also exist.

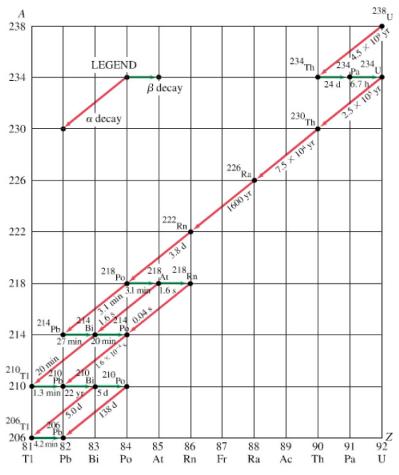


FIGURE 30-11 Decay series beginning with 238U. Nuclei in the series are specified by a dot representing A and Z values. Half-lives are given in seconds (s), minutes (min), hours (h), days (d), or years (yr). Note that a horizontal arrow represents β decay (A does not change), whereas a diagonal line represents α decay (A changes by 4, Z changes by 2).

Because of such decay series, certain radioactive elements are found in nature that otherwise would not be. For when the solar system acquired its present form about 5 billion years ago, it is believed that nearly all nuclides were formed (by the fusion process, Sections 31-3 and 33-2). Many isotopes with short half-lives decayed quickly and no longer are detected in nature today. But long-lived isotopes, such as $^{238}_{92}$ U with a half-life of 4.5×10^9 yr, still do exist in nature today. Indeed, about half of the original 238U still remains (assuming that the origin of the solar system was about 5×10^9 yr ago). We might expect, however, that radium (226/88Ra), with a half-life of 1600 yr, would long since have disappeared from the Earth. Indeed, the original 226Ra nuclei must by now have all decayed. However, because ²³⁸U decays (in several steps) to ²²⁶Ra, the supply